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Average titre volume = $(15.1 + 15.2)/2 = 15.15 \text{ cm}^3$. Number of moles of sulfuric acid used = $0.01515 \times 0.05 = 7.575 \times 10^{-4} \text{ mol}$ Number of moles of NaOH left in the 25.0 cm^3 hydrolysed solution = $2 \times 7.575 \times 10^{-4} = 1.515 \times 10^{-3} \text{ mol}$ Number of moles of NaOH left in 250.0 cm^3 of the hydrolysed solution = $10 \times 1.515 \times 10^{-3}$.

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Electron pairs = number of electrons on centre + number of bonded atoms/2. If we apply this to a molecule of ammonia NH_3 :- 5 outer electrons on the central nitrogen atom (electron arrangement 2, 5)+3hydrogen atoms bonded. $5+3(8)/2 = 4$ (3 bonded and 1 lone pair) Total number of electron pairs Shape.