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 the left. Show all your
 work in the space
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 actual yield of a
 reaction is 22 g and
 the theoretical yield is

25 g. Calculate the
 percentage yield. 2.
 6.0 mol of N₂ are
 mixed with 12.0 mol of
 H₂ according to the
 following equation: N₂(g)
 + 3H₂(g) → 2NH₃(g)

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 mixed with 12.0 mol of
 H₂ according to the
 following equation: N₂(g)
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Review Answers

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Stoichiometry

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ANSWER Answer the following questions in the space provided. 1. Given the following equation: $\text{C}_3\text{H}_4(\text{g}) + x \cdot \text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ?

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Stoichiometry
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Stoichiometry MIXED
REVIEW SHORT
ANSWER Answer the following questions in the space provided. 1. Given the following equation: $\text{C}_3\text{H}_4(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) +$

$2\text{H}_2\text{O}(\text{g})$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O 2:1 mol H 2O

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equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is
the value of the
coefficient x in this
equation? 40.07 g/mol
b.

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value of the coefficient
. x. in this equation? b.
What is the molar mass
of C_3H_4 ? c. How many
moles are in an 8.0 g
sample of C_3H_4 ? 2. a.
What is meant by .
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Stoichiometry b.
Theoretically, how
many moles of NH_3 will

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REVIEW.

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