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Ap Chemistry
Chapter 14 Answers
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 14 Answers - Zumdahl
 14.39 a. $\text{HClO}_4(\text{aq}) +$
 $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) +$
 $\text{ClO}_4^-(\text{aq})$. Only the
 forward reaction is
 indicated since HClO_4
 is a strong acid and is
 basically 100%
 dissociated in water.
 For acids, the
 dissociation reaction is

commonly written
 without water as a
 reactant. The common
 abbreviation for this

AP Chemistry **Chapter 14 Answers** **- Zumdahl 14**

AP Chemistry Chapter
 14 Answers - Zumdahl
 14.41 $\text{pH} = -\log[\text{H}^+]$;
 $\text{pOH} = -\log[\text{OH}^-]$; At 25
 $^\circ\text{C}$, $\text{pH} + \text{pOH} = 14.00$;
 For Exercise 13.37: a.
 $\text{pH} = -\log[\text{H}^+] = -$
 $\log(1.0 \times 10^{-7}) = 7.00$;
 $\text{pOH} = 14.00 - \text{pH} =$
 $14.00 - 7.00 = 7.00$ b.
 $\text{pH} = -\log(8.3 \times 10^{-16})$
 $= 15.08$; $\text{pOH} = 14.00$
 $- 15.08 = -1.08$ c. $\text{pH} =$
 $-\log(12) = -1.08$; pOH
 $= 14.08 - (-1.08) =$
 15.08 d.

AP Chemistry **Chapter 14 Answers** **- Zumdahl 14**

AP Chemistry Chapter
 7 Review Questions Ap
 chemistry chapter 14
 review questions

answers. Multiple-choice exercise. Use these values to assist you in answering some of the questions in this problems set Use the periodic table below to answer the questions in this Chapter Test. In the space provided, . . . Ap chemistry chapter 14 review questions answers.

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AP Chemistry Chapter
14. Chemical Kinetics -
1 - Change in time
Change in the
concentration of B
Average rate with
respect to B = Chapter
14. Chemical Kinetics
14.1 Factors that Affect
Reaction Rates •
Chemical kinetics =
the study of how fast
chemical reactions
occur. • Factors which

affect rates of
reactions:

Chapter 14. Chemical Kinetics

AP CHEMISTRY - MCHS
Page | 9 14.5

Temperature and Rate
(EOCQ'S 47 - 59 page
620) • Most reactions
speed up as
temperature increases.
• We can illustrate this
with chemiluminescent
Cyalume® light sticks.
• A chemiluminescent
reaction produces
light. • Two light sticks
are placed in water,
one at room
temperature and one
in ice.

Chapter 14. Chemical Kinetics

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Answer: Using Equation
14.14, we find $[\text{NO}_2] = 1.00 \times 10^{-3} \text{ M}$
Practice Exercise

Solution (continued)
18. Sample Exercise
14.9 Determining the
Half-Life of a First-
Order Reaction The
reaction of $\text{C}_4\text{H}_9\text{Cl}$
with water is a first-
order reaction.

**AP Chemistry
Chapter 14 Sample
Exercises**

Answer: carbon, by
 18.1°C Calculate the
work done for the
expansion of oxygen
gas from 2.5 liters to
3.7 liters against a
pressure of 720 torr.
Answer: $-1.14 \text{ liter}\cdot\text{atm}$
20.0 mL of pure water
at 285 K is mixed with
48 mL of water at 315
K. What is the final
temperature of the
mixture in kelvins?

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is an example ...
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10. Gases Chapter 10.
Gases. AP Chemistry
Chapter 10. Gases - 4 -
10.3 The Gas Laws To
describe a gas, you
need: Volume Pressure
Temperature (K)
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Language: English;
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