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The gas particles take up more volume

relative to the overall volume. The gas particles become hotter, so they increase the pressure along the container wall.

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Piersa, Amanda / Behavior of Gases

1. Gas particles are so small in relation to the distances between them that their individual volumes can be assumed to be insignificant. 2. No attractive or repulsive forces exist between gas particles. 3. Gas particles are in constant random motion, traveling in independent straight paths until they collide with each other or other objects.

The Behavior Of Gases Chapter 14 - ProProfs Quiz

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14.4 Gases: Mixtures and Movements

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