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Electron Configuration Test And Answer

In atomic physics and quantum chemistry, the electron configuration is the distribution of electrons of an atom or molecule (or other physical structure) in atomic or molecular orbitals. For example, the electron configuration of the neon atom is $1s^2 2s^2 2p^6$, using the notation explained below. Electronic configurations describe each electron as moving independently in an orbital, in an average field created by all other orbitals.

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It is important, therefore, to understand the arrangement of an atom's electrons. This 10-question multiple-choice chemistry practice test deals with the concepts of electronic structure, Hund's Rule, quantum numbers, and the Bohr atom. Answers to the questions appear at the end of the test. Question 1

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Electron Configuration Test With Answers

(d) Predict the charge of the most common titanium ion. Justify your answer. 7. Consider an atom of tin (Sn) in the ground state. (a) Write the electron configuration for a ground state atom of tin the noble gas . Use abbreviation. (b) Give a valid set of quantum numbers for the highest energy electron in a ground state atom of tin. 8.

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ground state electron configuration is $[\text{Kr}]5s^24d^5$. Since it is a transition metal, it takes less energy to remove electrons from the 4s shell because it is further away from the nucleus. Tc^{3+} would be $[\text{Kr}]4d^4$ because you remove 2 electrons from the s

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