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A Level OCR Chemistry Electrochemical Cells Questions ...

Electrochemical cells have two conductive electrodes, called the anode and the cathode. The anode is defined as the electrode where oxidation occurs. The cathode is the electrode where reduction takes place. Electrodes can be made from any sufficiently conductive materials, such as metals, semiconductors, graphite, and even conductive polymers.

Electrochemical Cells | Boundless Chemistry

Using E (cell) = E (red) - E (ox) E (cell) = -0.44 - 0 = -0.44V. This is negative and therefore the reaction will not happen, it is not spontaneous. However for the reaction: 2Fe 3+ + H 2 --> 2Fe 2+ + 2H +. E (cell) = +0.77 - 0.0 = +0.77. This reaction is spontaneous and therefore can happen.

Electrochemical cells - The Student Room

Electrochemical cells are constructed of various materials, such as the wire, the solutions themselves, and the containers. All of these materials cause the cell to have a property called internal...

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Practical 18 - Measuring the emf of an electrochemical cell (Required Practical 8) Click here to view some great books which can aid your learning For latest news check www.mwalimuluke.wordpress.com

Topic 13 - Electrochemistry - A-Level Chemistry

An electrochemical cell converts chemical energy into electrical energy using a redox reaction. Since, metals can be oxidised or reduced depending upon their chemical environment, then such an arrangement as shown below may be set-up. The Daniell cell, specifically, uses Zn (s) /Zn 2+ (aq) and Cu (s) /Cu 2+ (aq) reactions. Note that the two rods in the diagram are called electrodes.

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