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The Mole & Stoichiometry Study Guide (with answers) Solve each problem. SHOW ALL WORK FOR CREDIT! 1. How many molecules are in 4.85 mol sulfur dioxide? (2.92 x 10²⁴ molecules) 2. What is the mass in grams of 2.25 mol C₆H₆? (176 g) 3. What is the volume, in liters, of 0.895 mol of propane gas at STP? (20.0 L) 4.

Copy_of_Mole_Stoich_SG_with_ANSWERS_F20 - The Mole

...
Mole Calculations Study Guide 1 Answer Key Mole Calculations Study Guide Answer A mole of a particular substance is equal to the number of atoms in exactly 12 g of the carbon-12 isotope. Experiments established that number to be 6.022142 * 10²³ particles. Avogadro's number = NA = 6.022 * 10²³ items/mole.

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46 g carbon = 2.3 * 10²⁴ atoms. Consider the following: One mole of carbon is 12.01 g and contains 6.022 * 10²³ atoms. One mole of oxygen (O₂) is 32.00 g and contains 6.022 * 10²³ molecules. One mole of carbon dioxide (CO₂) is 44.01 g and contains 6.022 * 10²³ molecules.

Stoichiometry and the Mole Study Guide - Ms. Osawaru

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