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3. Another example: $7.45 \times 10^{-8} = 0.0000000745$ the "-8" tells you that there are 8 decimal places between the right side of the 7 and the end of the number.

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24 ChemQuest 8 Name: _____ Date: _____
 Hour: _____ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

24 ChemQuest 8 Name:

Hint: add your answer to question 1 and four times your answer to question 2. n\$ t ?(q\ = BZ e*eeAronS Verify that there are 32 electrons pictured in the drawing of CCl₄. 3es Find the sum of all the valence electrons for NF₃. (Add how many valence electrons one nitrogen atom has with the valence electrons for three fluorine atoms.) H: ?crl 2 Tte\}e.

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HW Keys - Roosevelt High School AP Chemistry 2017-18

Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S. Step 2: Criss cross to balance charges. Step 1: Write each ion and each charge. Positive first! Here's what you get: Add parentheses balance charges. Step 2. Criss cross to

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Hint: add your answer to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$ 4. Verify that there are 32 electrons pictured in the drawing of CCl_4 . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4

lines (each line represents two electrons). 5.

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